

1. What was Dalton's theory of the structure of the atom?
2. How did Thomson's cathode ray experiment contribute to the development of atomic theory?
3. According to Rutherford's model, all of an atom's positive charge is concentrated in its _____.
4. Draw and name the 5 models of atomic theory in chronological order. How did each model improve on the previous model?
5. In the black light lab, use the concepts of ground and excited states to explain how UV light caused your drawing and the skeleton to glow.
6. In the flame test lab, why did the substances produce light of different colors?

7. Complete the following chart:

Properties of Subatomic Particles

Particle	Charge	Location
Electron		
Proton		
Neutron		

8. Complete the table below by filling in the missing numbers.

Element	Atomic Number	Mass Number	Number of Protons	Number of Neutrons	Number of Electrons
Hydrogen					
Helium					
Carbon					
Fluorine					

9. How are the isotopes of an element different from one another?

10. Draw the Bohr model for a nitrogen atom.