

1. What is a **heterogeneous** mixture? Give an example.

2. Define **malleability**-

3. What is the difference between chemical and physical changes?

4. T or F A **compound** has the same properties as the elements that make it up.
Explain.

5. Distinguish between a **solution**, **suspension**, and a **colloid**.

6. How do mixtures differ from pure substances?

7. What does it mean if a substance has a high **reactivity**?

8. Define the following prefixes:

hetero-

homo-

di-

hydro-

9. What is a **precipitate**?

10. Compare and contrast a mixture and a compound.

11. When you describe a liquid as thick, are you saying that it has a low or high **viscosity**?

12. What allows a mixture to be separated by distillation?

13. Explain why the rusting of an iron bar decreases the strength of the bar.

16. Explain why silicon dioxide cannot be the only compound in the sample of sand shown in figure 5 on page 41.

17. How does the AMA use density to test the purity of methanol?

18. Using figure 12 on page 47, which of the substances are liquids at room temperature?