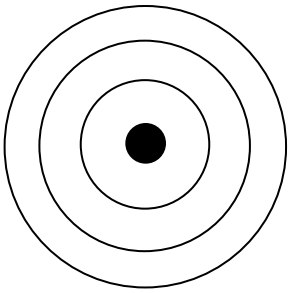
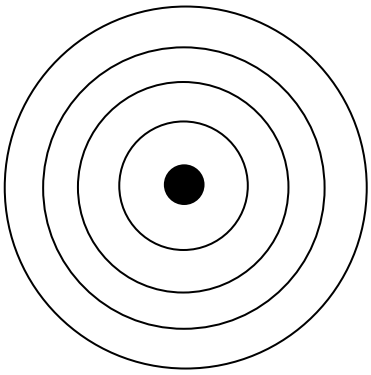
Name: _____
Date: _____
Period: _____

Bohr Model Practice

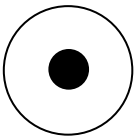
For each element, write the total number of electrons on the line. Then color the correct number of electrons for each orbit. Remember, fill the orbit closest to the nucleus first, but never exceed the number each orbit can hold. *Check the Periodic Table to find out how many electrons each element actually has*.



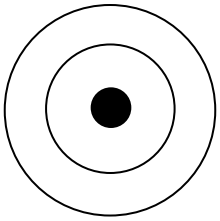
Sodium (Na) _____



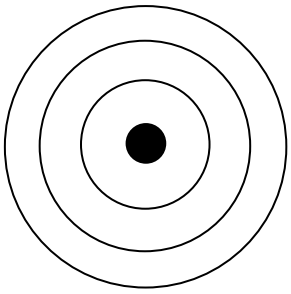
Potassium (K) _____



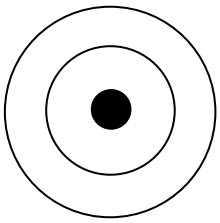
Hydrogen (H) _____



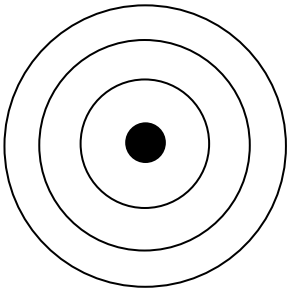
Carbon (C) _____

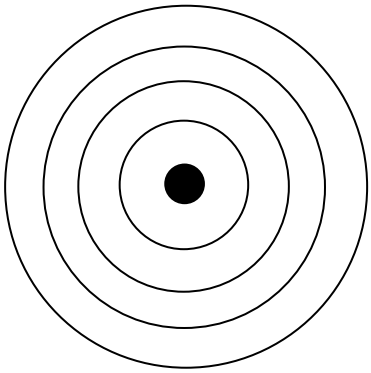


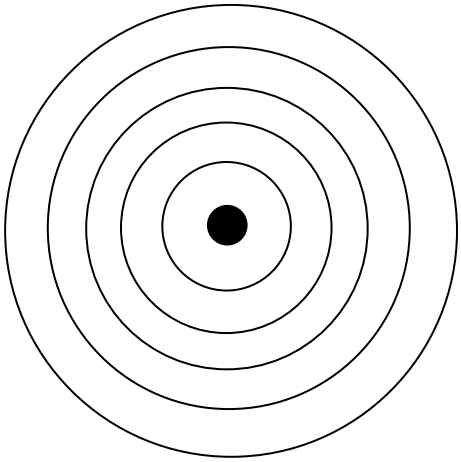
Silicon (Si) _____



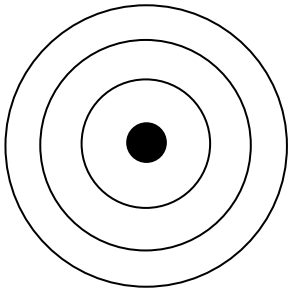
Oxygen (O) _____

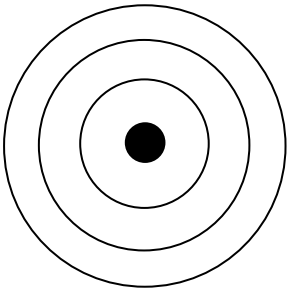


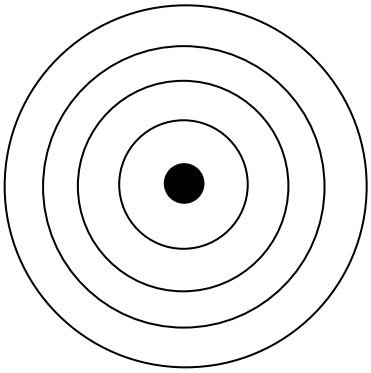




Chlorine (Cl) _____ Bromine (Br) _____ Iodine (I) _____







Argon (Ar) _____ Magnesium (Mg) _____ Calcium (Ca) _____

Now draw your own Bohr model diagrams for the following atoms:

Lithium (Li) _____ Sulfur (S) _____ Neon (Ne) _____